



Cambridge O Level

COMBINED SCIENCE

5129/21

Paper 2 Theory

October/November 2021

MARK SCHEME

Maximum Mark: 100

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge International will not enter into discussions about these mark schemes.

Cambridge International is publishing the mark schemes for the October/November 2021 series for most Cambridge IGCSE™, Cambridge International A and AS Level components and some Cambridge O Level components.

This document consists of **13** printed pages.

PUBLISHED**Generic Marking Principles**

These general marking principles must be applied by all examiners when marking candidate answers. They should be applied alongside the specific content of the mark scheme or generic level descriptors for a question. Each question paper and mark scheme will also comply with these marking principles.

GENERIC MARKING PRINCIPLE 1:

Marks must be awarded in line with:

- the specific content of the mark scheme or the generic level descriptors for the question
- the specific skills defined in the mark scheme or in the generic level descriptors for the question
- the standard of response required by a candidate as exemplified by the standardisation scripts.

GENERIC MARKING PRINCIPLE 2:

Marks awarded are always **whole marks** (not half marks, or other fractions).

GENERIC MARKING PRINCIPLE 3:

Marks must be awarded **positively**:

- marks are awarded for correct/valid answers, as defined in the mark scheme. However, credit is given for valid answers which go beyond the scope of the syllabus and mark scheme, referring to your Team Leader as appropriate
- marks are awarded when candidates clearly demonstrate what they know and can do
- marks are not deducted for errors
- marks are not deducted for omissions
- answers should only be judged on the quality of spelling, punctuation and grammar when these features are specifically assessed by the question as indicated by the mark scheme. The meaning, however, should be unambiguous.

GENERIC MARKING PRINCIPLE 4:

Rules must be applied consistently, e.g. in situations where candidates have not followed instructions or in the application of generic level descriptors.

GENERIC MARKING PRINCIPLE 5:

Marks should be awarded using the full range of marks defined in the mark scheme for the question (however; the use of the full mark range may be limited according to the quality of the candidate responses seen).

GENERIC MARKING PRINCIPLE 6:

Marks awarded are based solely on the requirements as defined in the mark scheme. Marks should not be awarded with grade thresholds or grade descriptors in mind.

Science-Specific Marking Principles

- | | |
|---|--|
| 1 | Examiners should consider the context and scientific use of any keywords when awarding marks. Although keywords may be present, marks should not be awarded if the keywords are used incorrectly. |
| 2 | The examiner should not choose between contradictory statements given in the same question part, and credit should not be awarded for any correct statement that is contradicted within the same question part. Wrong science that is irrelevant to the question should be ignored. |
| 3 | Although spellings do not have to be correct, spellings of syllabus terms must allow for clear and unambiguous separation from other syllabus terms with which they may be confused (e.g. ethane / ethene, glucagon / glycogen, refraction / reflection). |
| 4 | The error carried forward (ecf) principle should be applied, where appropriate. If an incorrect answer is subsequently used in a scientifically correct way, the candidate should be awarded these subsequent marking points. Further guidance will be included in the mark scheme where necessary and any exceptions to this general principle will be noted. |
| 5 | <p><u>'List rule' guidance</u></p> <p>For questions that require <i>n</i> responses (e.g. State two reasons ...):</p> <ul style="list-style-type: none"> • The response should be read as continuous prose, even when numbered answer spaces are provided. • Any response marked <i>ignore</i> in the mark scheme should not count towards <i>n</i>. • Incorrect responses should not be awarded credit but will still count towards <i>n</i>. • Read the entire response to check for any responses that contradict those that would otherwise be credited. Credit should not be awarded for any responses that are contradicted within the rest of the response. Where two responses contradict one another, this should be treated as a single incorrect response. • Non-contradictory responses after the first <i>n</i> responses may be ignored even if they include incorrect science. |

6 Calculation specific guidance

Correct answers to calculations should be given full credit even if there is no working or incorrect working, **unless** the question states 'show your working'.

For questions in which the number of significant figures required is not stated, credit should be awarded for correct answers when rounded by the examiner to the number of significant figures given in the mark scheme. This may not apply to measured values.

For answers given in standard form (e.g. $a \times 10^n$) in which the convention of restricting the value of the coefficient (a) to a value between 1 and 10 is not followed, credit may still be awarded if the answer can be converted to the answer given in the mark scheme.

Unless a separate mark is given for a unit, a missing or incorrect unit will normally mean that the final calculation mark is not awarded. Exceptions to this general principle will be noted in the mark scheme.

7 Guidance for chemical equations

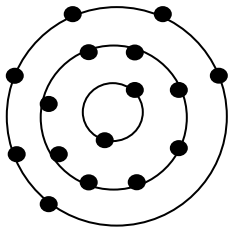
Multiples / fractions of coefficients used in chemical equations are acceptable unless stated otherwise in the mark scheme.

State symbols given in an equation should be ignored unless asked for in the question or stated otherwise in the mark scheme.

Question	Answer	Marks
1	dissolve ; solvent ; filtered ; filtrate ; crystallises ;	5

Question	Answer	Marks
2(a)(i)	6 (m/ s) ;	1
2(a)(ii)	4 m / s for 3 s or 4 m / s from 4 to 7 s or constant speed for 3 s ; non-constant speed increase from 4 m / s to 12 m / s or from 7 s to 12 s or for 5 s ; 12 m / s to 0 m/ s in 3 s or decreases to 0 m / s in 3 s ;	3

Question	Answer	Marks
3	<i>similarity:</i> any one from : <ul style="list-style-type: none"> • movement is from high concentration to low concentration ; • movement caused by random movement of particles / does not require energy ; <i>difference:</i> any two from : <ul style="list-style-type: none"> • osmosis applies to water only ; • osmosis takes place though a partially permeable membrane only ; • diffusion can occur in gases ; 	3

Question	Answer	Marks
4(a)		1
4(b)	covalent ;	1
4(c)(i)	64 ;	1
4(c)(ii)	128 36 ;; 6.4 ;	3
4(d)	5–6.5 ;	1

Question	Answer	Marks
5(a)	any four from : <ul style="list-style-type: none"> • use of stopwatch / timer ; • displace and release bob ; • stop (timer) when bob returns to starting position ; • time for more than one swing ; • divide total time by number of swings / calculate the average time ; 	4

Question	Answer	Marks
5(b)	<p> diameter of bob micrometer length of thread metre rule with cm divisions width of thread vernier caliper </p>	2

Question	Answer	Marks
6(a)	3 electron shells ;	1
6(b)	group number is number of electrons in outer shell ;	1
6(c)	alkali metals ;	1
6(d)(i)	metallic to non-metallic ;	1
6(d)(ii)	sodium oxide – basic ; phosphorus oxide – acidic ;	2
6(e)	it is stable ;	1

Question	Answer	Marks
7(a)(i)	3100 (mm ²) ;	1
7(a)(ii)	10 (arbitrary units) ;	1
7(a)(iii)	in C , as the light intensity decreases, the surface area (of leaves) increases ; in D , as the light intensity increases from 10–50 au, the surface area (of leaves) increases ; in D as the light intensity increases from 50–100 au, the surface area of the leaves decreases;	3
7(b)	(average leaf size changes as light intensity changes) to absorb enough / sufficient light ; for photosynthesis ;	2

Question	Answer	Marks
8(a)	clockwise moment ; = / balances anticlockwise moment ;	2
8(b)	(moment =) $f \times d$ or $5.0 \times 250\,000$; 1 250 000 (Nm) ;	2
8(c)	$20 \times 125\,000 \div 250\,000$ or 20×0.5 ; 10 (m) ;	2

Question	Answer	Marks				
9(a)	<table border="1"> <tr> <td>process</td> <td>letter from Fig. 4.1</td> </tr> <tr> <td>absorption digestion egestion ingestion</td> <td>J / L / M ; H / J / L ; N ; H ;</td> </tr> </table>	process	letter from Fig. 4.1	absorption digestion egestion ingestion	J / L / M ; H / J / L ; N ; H ;	4
process	letter from Fig. 4.1					
absorption digestion egestion ingestion	J / L / M ; H / J / L ; N ; H ;					
9(b)	<i>enzyme name</i> : amylase ; <i>acting on</i> : starch / carbohydrate ;	2				

Question	Answer	Marks
10(a)	20–21% ; 78–79% ;	2
10(b)	any one from <ul style="list-style-type: none"> • making ammonia ; • making fertilisers ; • making nitric acid ; 	1
10(c)	acid rain ;	1
10(d)(i)	$4\text{Li} + \text{O}_2 \longrightarrow 2\text{Li}_2\text{O}$;	1
10(d)(ii)	Li_3N ;	1

Question	Answer	Marks
11(a)	arrow labelled F at 0 ; arrow labelled B at 100 ;	2
11(b)(i)	26 (°C)	1
11(b)(ii)	thermal / heat energy is transferred from water to liquid in thermometer ; (so that) liquid inside thermometer expands ;	2
11(c)	any one from: <ul style="list-style-type: none"> • narrower central tube (bore) • larger bulb 	1

Question	Answer	Marks																								
12	<table border="1" style="width: 100%; border-collapse: collapse;"> <thead> <tr> <th></th> <th>aerobic only</th> <th>anaerobic only</th> <th>both</th> </tr> </thead> <tbody> <tr> <td>releases CO₂</td> <td style="text-align: center;">√</td> <td></td> <td></td> </tr> <tr> <td>releases energy</td> <td></td> <td></td> <td style="text-align: center;">√</td> </tr> <tr> <td>use glucose</td> <td></td> <td></td> <td style="text-align: center;">√</td> </tr> <tr> <td>use O₂</td> <td style="text-align: center;">√</td> <td></td> <td></td> </tr> <tr> <td>produce H₂O</td> <td style="text-align: center;">√</td> <td></td> <td></td> </tr> </tbody> </table>		aerobic only	anaerobic only	both	releases CO ₂	√			releases energy			√	use glucose			√	use O ₂	√			produce H ₂ O	√			5
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Question	Answer	Marks
13(a)	A = hydrogen ; B = water ; C = poly(ethene) ;	3
13(b)(i)	contains a <u>carbon to carbon</u> double bond ;	1
13(b)(ii)	(aqueous) bromine ;	1
13(c)	exothermic ;	1
13(d)	$ \begin{array}{c} \text{H} \quad \text{H} \\ \quad \\ \text{H} - \text{C} - \text{C} - \text{H} \\ \quad \\ \text{H} \quad \text{H} \end{array} $	1

Question	Answer	Marks
14(a)(i)	amplitude marked with an arrow ;	1
14(a)(ii)	$589 (\times 10^{-9} \text{ m})$;	1
14(b)	(P) ultraviolet ; (Q) infrared ;	2
14(c)	(f =) v / λ or $3 \times 10^8 \div 589 \times 10^{-9}$; $5.09(338) \times 10^{14}$; $5.1 \times 10^{14} (\text{Hz})$;	3

Question	Answer	Marks
15	<pre> graph LR A[anther] --> B[produces pollen] C[carpel] --> D[produces seeds] E[root] --> F[absorbs water] G[xylem] --> H[transports water] </pre>	4

Question	Answer	Marks
16	most reactive Y Z X least reactive W ;;	2

Question	Answer	Marks
17(a)	transformer ;	1
17(b)	magnetic field ; (magnetic field) changes direction/alternates or (magnetic field) cuts secondary coil ; idea of different number of coils/turns (on primary compared with secondary) ;	3

Question	Answer	Marks
18	(blood) capillary ;	1
	<i>Vein linked to:</i> contains valves ; have a wider lumen than arteries ; usually carry deoxygenated blood ;	3

Question	Answer	Marks
19	sperm ducts ; diaphragm ; condom ; (contraceptive) pill ;	4